

chapter 6 thermochemistry pdf

Thermochemistry (Chapter 6) General Concepts and Terminology 1. Kinetic and Potential Energy -- review 2. Chemical Energy: Potential energy associated with chemical bonds: ... 6. Hess' Law of Heat Summation calculate ΔH° for a reaction from tabulated ΔH°_f values

Thermochemistry (Chapter 6) General Concepts and Terminology

Chapter 6: Thermochemistry Thermochemistry: energy considerations associated with chemical and physical change Energy: the capacity of a system to do work or produce heat potential energy ΔE energy of position; stored energy kinetic energy ΔE energy of motion kinetic energy = $\frac{1}{2}mv^2$ kinetic energy = $3nRT$

Chapter 6: Thermochemistry Thermochemistry: energy

CHAPTER 6: THERMOCHEMISTRY 163 Now, we substitute P and ΔV into Equation (6.3) of the text to solve for w. $w = \Delta P \Delta V = \Delta P (1.0 \text{ atm})(31 \text{ L}) = \Delta P 31 \text{ L} \cdot \text{atm}$ The problems asks for the work done in units of joules. The following conversion factor can be obtained

CHAPTER 6 THERMOCHEMISTRY - Oregon State University

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Thermochemistry: study of energy changes in a system during chemical or physical changes. Internal Energy (U): total energy of a system. Heat (q): quantity of energy transferred between a system and its surroundings. (Remember a system does not ... Chapter 6: Thermochemistry Author: Sapna Gupta Created Date:

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CHAPTER 6 THERMOCHEMISTRY 131 b.)E = 82 + 47 = 129 kJ c.)E = 47 + 0 = 47 kJ d. When the surroundings deliver work to the system, $w > 0$. This is the case for a and b.

CHAPTER SIX THERMOCHEMISTRY - bremertonschools.org

Chapter 6 Thermochemistry Concept Check 6.1 A solar-powered water pump has photovoltaic cells on protruding top panels. These cells collect energy from sunlight, storing it momentarily in a battery, which later runs an electric motor that pumps water up to a storage tank on a hill.

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CHAPTER 6 Thermochemistry 1. A gas absorbs 0.0 J of heat and then performs 15.2 J of work. The change in internal energy of the gas is a) b) c) d) e)

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6-1 CHAPTER 6 THERMOCHEMISTRY: ENERGY FLOW AND CHEMICAL CHANGE 6.1 The sign of the energy transfer is defined from the perspective of the system. Entering the system is positive, and leaving the system is negative. 6.2 No, an increase in temperature means that heat has been transferred to the surroundings, which makes q positive.

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